

Tetraaquabis(5-hydroxynicotinato- κ N)-cadmium(II)

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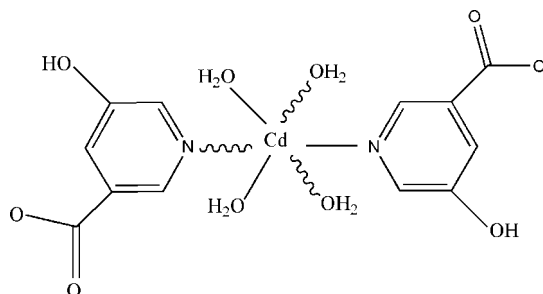
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Key indicators: single-crystal X-ray study; $T = 296$ K; mean $\sigma(\text{C}-\text{C}) = 0.002$ Å; R factor = 0.016; wR factor = 0.042; data-to-parameter ratio = 13.4.

The title compound, $[\text{Cd}(\text{C}_6\text{H}_4\text{NO}_3)_2(\text{H}_2\text{O})_4]$, was obtained by the reaction of cadmium chloride with 5-hydroxynicotinic acid. The Cd^{II} atom is located on an inversion centre and is coordinated by two N atoms from two 5-hydroxynicotinic acid ligands and four water molecules in a distorted octahedral geometry. The structure is stabilized by intermolecular $\text{O}-\text{H}\cdots\text{O}$ hydrogen bonds, forming a three-dimensional network.

Related literature

For cadmium compounds and their photoluminescent properties, see: He *et al.* (2008); Kang *et al.* (2007); Zhang *et al.* (2006); Zora *et al.* (2006).



Experimental

Crystal data

$[\text{Cd}(\text{C}_6\text{H}_4\text{NO}_3)_2(\text{H}_2\text{O})_4]$
 $M_r = 460.68$

Triclinic, $P\bar{1}$
 $a = 7.2190$ (1) Å

$b = 7.2510$ (1) Å
 $c = 8.9260$ (1) Å
 $\alpha = 70.377$ (1)°
 $\beta = 68.154$ (1)°
 $\gamma = 65.7170$ (10)°
 $V = 385.97$ (1) Å³

$Z = 1$
Mo $K\alpha$ radiation
 $\mu = 1.48$ mm⁻¹
 $T = 296$ (2) K
 $0.27 \times 0.17 \times 0.07$ mm

Data collection

Bruker APEXII diffractometer
Absorption correction: multi-scan
(*SADABS*; Sheldrick, 1996)
 $T_{\text{min}} = 0.667$, $T_{\text{max}} = 0.903$

6067 measured reflections
1759 independent reflections
1754 reflections with $I > 2\sigma(I)$
 $R_{\text{int}} = 0.017$

Refinement

$R[F^2 > 2\sigma(F^2)] = 0.016$
 $wR(F^2) = 0.042$
 $S = 1.09$
1759 reflections
131 parameters
7 restraints

H atoms treated by a mixture of independent and constrained refinement
 $\Delta\rho_{\text{max}} = 0.36$ e Å⁻³
 $\Delta\rho_{\text{min}} = -0.34$ e Å⁻³

Table 1

Hydrogen-bond geometry (Å, °).

$D-\text{H}\cdots A$	$D-\text{H}$	$\text{H}\cdots A$	$D\cdots A$	$D-\text{H}\cdots A$
$\text{O1W}-\text{H1WA}\cdots\text{O2}^{\text{i}}$	0.81 (2)	1.94 (2)	2.742 (2)	171 (3)
$\text{O1W}-\text{H1WB}\cdots\text{O3}^{\text{ii}}$	0.79 (2)	2.20 (2)	2.973 (2)	164 (3)
$\text{O2W}-\text{H2WA}\cdots\text{O1}^{\text{iii}}$	0.82 (2)	1.87 (2)	2.656 (2)	160 (3)
$\text{O2W}-\text{H2WB}\cdots\text{O2}^{\text{iv}}$	0.82 (2)	1.93 (2)	2.735 (2)	165 (3)
$\text{O3}-\text{H3}\cdots\text{O1}^{\text{v}}$	0.83 (2)	1.88 (2)	2.664 (2)	157 (3)

Symmetry codes: (i) $-x+1, -y+1, -z+1$; (ii) $-x, -y+1, -z+2$; (iii) $-x+2, -y+1, -z+1$; (iv) $x, y-1, z+1$; (v) $x-1, y, z$.

Data collection: *SMART* (Bruker, 2004); cell refinement: *S SAINT* (Bruker, 2004); data reduction: *XPREP* (Bruker, 2004); program(s) used to solve structure: *SHELXTL* (Sheldrick, 2008); program(s) used to refine structure: *SHELXTL*; molecular graphics: *SHELXTL*; software used to prepare material for publication: *SHELXTL*.

Supplementary data and figures for this paper are available from the IUCr electronic archives (Reference: AT2660).

References

- Bruker (2004). *S SAINT*, *SMART* and *XPREP*. Bruker AXS Inc., Madison, Wisconsin, USA.
- He, Y. H., Feng, Y. L., Lan, Y. Z. & Wen, Y. H. (2008). *Cryst. Growth Des.* **8**, 3586–3594.
- Kang, Y., Zhang, J., Qin, Y. Y., Li, Z. J. & Yao, Y. G. (2007). *J. Mol. Struct.* **784**, 98–108.
- Sheldrick, G. M. (1996). *SADABS*. University of Göttingen, Germany.
- Sheldrick, G. M. (2008). *Acta Cryst.* **A64**, 112–122.
- Zhang, L. Y., Zhang, J. P., Lin, Y. Y. & Chen, X. M. (2006). *Cryst. Growth Des.* **6**, 1684–1689.
- Zora, P., Gordana, P., Marijana, V. & Drazen, V. T. (2006). *Polyhedron*, **25**, 2353–2362.

supplementary materials

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Comment

There is intense research on the synthesis of the cadmium metal compounds for their interesting photoluminescent properties. A large number of these compounds have been synthesized (He *et al.*, 2008; Zora *et al.*, 2006; Kang *et al.*, 2007; Zhang *et al.*, 2006).

As illustrated in Fig. 1, the Cd(II) atom is coordinated by two nitrogen atoms from two 5-hydroxynicotinic acid ligands and four water molecules. Four coordinated atoms of O1W, O2W, O1WA and O2WA constitute the base of the octahedral, whereas N1 and N1A atoms occupy the apical position. The intermolecular hydrogen bonds play an important role in the formation of the three-dimensional network. As shown in Fig. 2, the intermolecular O—H \cdots O hydrogen bonds link the neighboring molecules to a three-dimensional network.

Experimental

A mixture of 0.5 mmol 5-hydroxynicotinic acid and 0.5 mmol of cadmium chloride in 10 ml distilled water was stirred for 30 min at 323 K, then the reaction mixture was filtered and well shaped colourless crystals of the title compound was obtained from the mother liquor by slow evaporation at room temperature for several days.

Refinement

The H atoms bonded to C atoms were positioned geometrically [aromatic C—H = 0.93 Å and aliphatic C—H = 0.97 Å, $U_{\text{iso}}(\text{H}) = 1.2U_{\text{eq}}(\text{C})$]. The H atoms bonded to O atoms were located in a difference Fourier maps and refined with O—H distance restraints of 0.85 and $U_{\text{iso}}(\text{H}) = 1.5U_{\text{eq}}(\text{O})$.

Figures

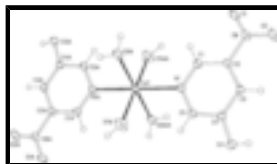


Fig. 1. A view of the molecule of (I), showing the atom-labelling scheme, displacement ellipsoids are shown at the 30% probability level. [Symmetry code: (A) $-x + 1, -y + 1, -z + 2$].



Fig. 2. A view of the three dimensional framework of the title compound. The O—H \cdots O interactions are depicted by dashed lines.

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Crystal data

[Cd(C ₆ H ₄ NO ₃) ₂ (H ₂ O) ₄]	$Z = 1$
$M_r = 460.68$	$F_{000} = 230$
Triclinic, $P\bar{1}$	$D_x = 1.982 \text{ Mg m}^{-3}$
Hall symbol: -P 1	Mo $K\alpha$ radiation
$a = 7.21900 (10) \text{ \AA}$	$\lambda = 0.71073 \text{ \AA}$
$b = 7.25100 (10) \text{ \AA}$	Cell parameters from 5615 reflections
$c = 8.92600 (10) \text{ \AA}$	$\theta = 2.5\text{--}27.5^\circ$
$\alpha = 70.3770 (10)^\circ$	$\mu = 1.48 \text{ mm}^{-1}$
$\beta = 68.1540 (10)^\circ$	$T = 296 (2) \text{ K}$
$\gamma = 65.7170 (10)^\circ$	Sheet, colourless
$V = 385.972 (9) \text{ \AA}^3$	$0.27 \times 0.17 \times 0.07 \text{ mm}$

Data collection

Bruker APEXII diffractometer	1759 independent reflections
Radiation source: fine-focus sealed tube	1754 reflections with $I > 2\sigma(I)$
Monochromator: graphite	$R_{\text{int}} = 0.017$
$T = 296(2) \text{ K}$	$\theta_{\text{max}} = 27.5^\circ$
ω scans	$\theta_{\text{min}} = 2.5^\circ$
Absorption correction: multi-scan (SADABS; Sheldrick, 1996)	$h = -9 \rightarrow 9$
$T_{\text{min}} = 0.667, T_{\text{max}} = 0.903$	$k = -9 \rightarrow 9$
6067 measured reflections	$l = -11 \rightarrow 11$

Refinement

Refinement on F^2	Secondary atom site location: difference Fourier map
Least-squares matrix: full	Hydrogen site location: inferred from neighbouring sites
$R[F^2 > 2\sigma(F^2)] = 0.016$	H atoms treated by a mixture of independent and constrained refinement
$wR(F^2) = 0.042$	$w = 1/[\sigma^2(F_o^2) + (0.0249P)^2 + 0.1253P]$
$S = 1.09$	where $P = (F_o^2 + 2F_c^2)/3$
1759 reflections	$(\Delta/\sigma)_{\text{max}} < 0.001$
131 parameters	$\Delta\rho_{\text{max}} = 0.36 \text{ e \AA}^{-3}$
7 restraints	$\Delta\rho_{\text{min}} = -0.33 \text{ e \AA}^{-3}$
Primary atom site location: structure-invariant direct methods	Extinction correction: none

Special details

Geometry. All e.s.d.'s (except the e.s.d. in the dihedral angle between two l.s. planes) are estimated using the full covariance matrix. The cell e.s.d.'s are taken into account individually in the estimation of e.s.d.'s in distances, angles and torsion angles; correlations between e.s.d.'s in cell parameters are only used when they are defined by crystal symmetry. An approximate (isotropic) treatment of cell e.s.d.'s is used for estimating e.s.d.'s involving l.s. planes.

Refinement. Refinement of F^2 against ALL reflections. The weighted R -factor wR and goodness of fit S are based on F^2 , conventional R -factors R are based on F , with F set to zero for negative F^2 . The threshold expression of $F^2 > \sigma(F^2)$ is used only for calculating R -factors(gt) *etc.* and is not relevant to the choice of reflections for refinement. R -factors based on F^2 are statistically about twice as large as those based on F , and R -factors based on ALL data will be even larger.

Fractional atomic coordinates and isotropic or equivalent isotropic displacement parameters (\AA^2)

	<i>x</i>	<i>y</i>	<i>z</i>	$U_{\text{iso}}^*/U_{\text{eq}}$
Cd1	0.5000	0.5000	1.0000	0.02626 (6)
O1	0.8938 (2)	0.7990 (2)	0.32463 (16)	0.0427 (3)
O1W	0.2684 (3)	0.3168 (3)	1.07212 (17)	0.0482 (4)
H1WA	0.287 (5)	0.249 (4)	1.009 (3)	0.072*
H1WB	0.209 (5)	0.270 (4)	1.163 (2)	0.072*
O2	0.6901 (2)	0.8709 (2)	0.16430 (15)	0.0429 (3)
O2W	0.7595 (2)	0.2074 (2)	0.92481 (18)	0.0478 (4)
H2WA	0.881 (3)	0.189 (4)	0.865 (3)	0.072*
H2WB	0.762 (5)	0.098 (3)	0.994 (3)	0.072*
O3	0.0066 (2)	0.7675 (2)	0.58671 (17)	0.0400 (3)
H3	0.002 (4)	0.785 (4)	0.491 (2)	0.053 (7)*
N1	0.4459 (2)	0.6297 (2)	0.74305 (16)	0.0264 (3)
C1	0.5945 (2)	0.6740 (2)	0.60554 (19)	0.0267 (3)
H1A	0.7280	0.6483	0.6138	0.032*
C2	0.5560 (2)	0.7568 (2)	0.45137 (18)	0.0245 (3)
C3	0.3595 (3)	0.7885 (2)	0.43778 (19)	0.0263 (3)
H3A	0.3310	0.8402	0.3354	0.032*
C4	0.2061 (2)	0.7415 (2)	0.5802 (2)	0.0270 (3)
C5	0.2549 (2)	0.6655 (3)	0.73035 (19)	0.0276 (3)
H5A	0.1507	0.6383	0.8260	0.033*
C6	0.7273 (3)	0.8123 (2)	0.30121 (19)	0.0288 (3)

Atomic displacement parameters (\AA^2)

	U^{11}	U^{22}	U^{33}	U^{12}	U^{13}	U^{23}
Cd1	0.02499 (9)	0.03841 (10)	0.01604 (8)	-0.01493 (7)	-0.00552 (6)	-0.00099 (6)
O1	0.0313 (6)	0.0688 (9)	0.0266 (6)	-0.0259 (6)	-0.0010 (5)	-0.0044 (6)
O1W	0.0589 (9)	0.0749 (10)	0.0276 (6)	-0.0484 (8)	-0.0017 (6)	-0.0091 (7)
O2	0.0587 (8)	0.0592 (8)	0.0199 (6)	-0.0385 (7)	-0.0088 (5)	0.0028 (5)
O2W	0.0368 (7)	0.0424 (7)	0.0347 (7)	-0.0075 (6)	0.0072 (6)	0.0013 (6)
O3	0.0285 (6)	0.0620 (8)	0.0326 (7)	-0.0220 (6)	-0.0135 (5)	0.0007 (6)
N1	0.0258 (6)	0.0362 (7)	0.0187 (6)	-0.0141 (5)	-0.0068 (5)	-0.0018 (5)

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C1	0.0252 (7)	0.0376 (8)	0.0208 (7)	-0.0155 (6)	-0.0064 (6)	-0.0037 (6)
C2	0.0278 (7)	0.0271 (7)	0.0195 (7)	-0.0128 (6)	-0.0046 (6)	-0.0033 (5)
C3	0.0312 (7)	0.0288 (7)	0.0207 (7)	-0.0120 (6)	-0.0109 (6)	-0.0005 (5)
C4	0.0249 (7)	0.0305 (7)	0.0281 (7)	-0.0112 (6)	-0.0107 (6)	-0.0026 (6)
C5	0.0254 (7)	0.0358 (8)	0.0219 (7)	-0.0143 (6)	-0.0053 (6)	-0.0023 (6)
C6	0.0340 (8)	0.0319 (8)	0.0206 (7)	-0.0166 (6)	-0.0025 (6)	-0.0038 (6)

Geometric parameters (\AA , $^\circ$)

Cd1—O2W	2.2830 (14)	O3—C4	1.3543 (19)
Cd1—O2W ⁱ	2.2830 (14)	O3—H3	0.830 (17)
Cd1—N1 ⁱ	2.2831 (13)	N1—C5	1.335 (2)
Cd1—N1	2.2831 (13)	N1—C1	1.3411 (19)
Cd1—O1W	2.3291 (13)	C1—C2	1.387 (2)
Cd1—O1W ⁱ	2.3291 (13)	C1—H1A	0.9300
O1—C6	1.255 (2)	C2—C3	1.385 (2)
O1W—H1WA	0.809 (17)	C2—C6	1.517 (2)
O1W—H1WB	0.794 (17)	C3—C4	1.389 (2)
O2—C6	1.244 (2)	C3—H3A	0.9300
O2W—H2WA	0.823 (17)	C4—C5	1.386 (2)
O2W—H2WB	0.822 (17)	C5—H5A	0.9300
O2W—Cd1—O2W ⁱ	180.0	C5—N1—C1	118.64 (13)
O2W—Cd1—N1 ⁱ	87.79 (5)	C5—N1—Cd1	117.86 (10)
O2W ⁱ —Cd1—N1 ⁱ	92.21 (5)	C1—N1—Cd1	123.49 (10)
O2W—Cd1—N1	92.21 (5)	N1—C1—C2	122.29 (14)
O2W ⁱ —Cd1—N1	87.79 (5)	N1—C1—H1A	118.9
N1 ⁱ —Cd1—N1	180.000 (1)	C2—C1—H1A	118.9
O2W—Cd1—O1W	85.72 (6)	C3—C2—C1	118.97 (14)
O2W ⁱ —Cd1—O1W	94.28 (6)	C3—C2—C6	121.06 (14)
N1 ⁱ —Cd1—O1W	90.57 (5)	C1—C2—C6	119.96 (14)
N1—Cd1—O1W	89.43 (5)	C2—C3—C4	118.68 (14)
O2W—Cd1—O1W ⁱ	94.28 (6)	C2—C3—H3A	120.7
O2W ⁱ —Cd1—O1W ⁱ	85.72 (6)	C4—C3—H3A	120.7
N1 ⁱ —Cd1—O1W ⁱ	89.43 (5)	O3—C4—C5	115.78 (14)
N1—Cd1—O1W ⁱ	90.57 (5)	O3—C4—C3	125.40 (14)
O1W—Cd1—O1W ⁱ	180.0	C5—C4—C3	118.81 (14)
Cd1—O1W—H1WA	117 (2)	N1—C5—C4	122.55 (14)
Cd1—O1W—H1WB	127 (2)	N1—C5—H5A	118.7
H1WA—O1W—H1WB	110 (2)	C4—C5—H5A	118.7
Cd1—O2W—H2WA	131 (2)	O2—C6—O1	125.02 (15)
Cd1—O2W—H2WB	117 (2)	O2—C6—C2	117.26 (15)
H2WA—O2W—H2WB	105 (2)	O1—C6—C2	117.71 (14)
C4—O3—H3	108.2 (19)		
O2W—Cd1—N1—C5	119.15 (12)	C1—C2—C3—C4	1.8 (2)
O2W ⁱ —Cd1—N1—C5	-60.85 (12)	C6—C2—C3—C4	-177.83 (14)

O1W—Cd1—N1—C5	33.45 (13)	C2—C3—C4—O3	178.81 (15)
O1W ⁱ —Cd1—N1—C5	-146.55 (13)	C2—C3—C4—C5	0.2 (2)
O2W—Cd1—N1—C1	-61.78 (13)	C1—N1—C5—C4	1.7 (2)
O2W ⁱ —Cd1—N1—C1	118.22 (13)	Cd1—N1—C5—C4	-179.16 (12)
O1W—Cd1—N1—C1	-147.47 (13)	O3—C4—C5—N1	179.22 (15)
O1W ⁱ —Cd1—N1—C1	32.53 (13)	C3—C4—C5—N1	-2.1 (2)
C5—N1—C1—C2	0.5 (2)	C3—C2—C6—O2	-5.8 (2)
Cd1—N1—C1—C2	-178.62 (11)	C1—C2—C6—O2	174.61 (15)
N1—C1—C2—C3	-2.2 (2)	C3—C2—C6—O1	173.06 (16)
N1—C1—C2—C6	177.42 (14)	C1—C2—C6—O1	-6.6 (2)

Symmetry codes: (i) $-x+1, -y+1, -z+2$.

Hydrogen-bond geometry (Å, °)

<i>D</i> —H... <i>A</i>	<i>D</i> —H	H... <i>A</i>	<i>D</i> ... <i>A</i>	<i>D</i> —H... <i>A</i>
O1W—H1WA...O2 ⁱⁱ	0.809 (17)	1.941 (17)	2.742 (2)	171 (3)
O1W—H1WB...O3 ⁱⁱⁱ	0.794 (17)	2.201 (18)	2.9728 (19)	164 (3)
O2W—H2WA...O1 ^{iv}	0.823 (17)	1.867 (18)	2.6556 (18)	160 (3)
O2W—H2WB...O2 ^v	0.822 (17)	1.933 (17)	2.7349 (19)	165 (3)
O3—H3...O1 ^{vi}	0.830 (17)	1.878 (19)	2.6637 (19)	157 (3)

Symmetry codes: (ii) $-x+1, -y+1, -z+1$; (iii) $-x, -y+1, -z+2$; (iv) $-x+2, -y+1, -z+1$; (v) $x, y-1, z+1$; (vi) $x-1, y, z$.

Fig. 1

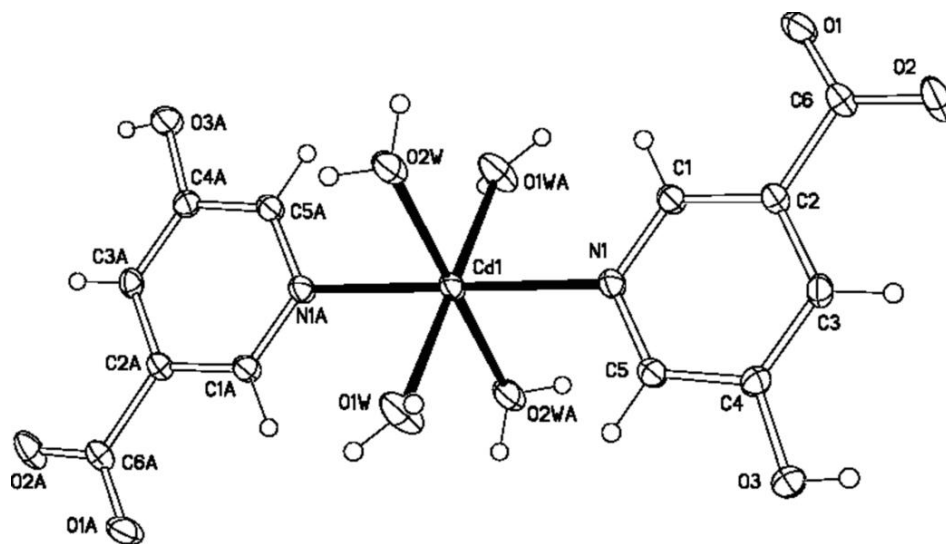


Fig. 2

